Name	

1. You must prepare 250.0 mL of 1.25 M KNO₃ solution using 3.00 M KNO₃ stock solution. How many milliliters of stock solution should you use?

2. How many grams of Ca(NO₃)₂ are needed to make 3.5 L of a 1.2 M solution?

3. Balance the equation $Fe + H_2SO_4 \rightarrow Fe_2(SO_4)_3 + H_2$ How many grams of iron (III) sulfate will be produced if 1.2 L of a 2.5 M solution of H_2SO_4 is reacted completely?

4.	66.4 g of Calcium Carbonate, CaCO ₃ are dissolved in 150.0 g of solution. What is the percent by mass?
5.	Calculate the molality of a 2.3 M solution of copper II sulfate, $CuSO_4$ in water, if the solution has a density of 1.19 g/mL.
6.	Find the boiling point of the 1.45 m solution of a water solution of zinc II chloride.